

Group 7 Reactivity

Statement	Could it be true?	Is the reasoning good?	Can you improve the statement?
Chlorine and fluorine will react in the same way because they are in the same group			
Chlorine is more reactive because it has more electrons			
Fluorine will be more reactive as it has fewer electrons to lose			
When chlorine and fluorine react they gain an electron to form negative ions			
Chlorine and fluorine react in the same way because they are the same atom with different numbers of electrons			
Fluorine is more reactive because it needs less energy to attract an electron because the outer shell is closer			
Fluorine will be more reactive because it has fewer electrons in its second shell			
Chlorine is more reactive because it has more protons to attract the extra electron			
Fluorine is more reactive than chlorine because when it gains an electron, the electron is closer to the nucleus so it is more strongly attracted			