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| --- | --- | --- | --- |
| Statement | Could it be true? | Is the reasoning good? | Can you improve the statement? |
| Chlorine and fluorine will react in the same way because they are in the same group |  |  |  |
| Chlorine is more reactive because it has more electrons to lose |  |  |  |
| Gaining an electron takes energy |  |  |  |
| When chlorine and fluorine react they gain an electron to form negative ions |  |  |  |
| Chlorine and fluorine react in the same way because they are the same atom with different numbers of electrons |  |  |  |
| Fluorine is more reactive because it needs less energy to attract an electron because the outer shell is closer |  |  |  |
| Fluorine will be more reactive because it has fewer electrons in its second shell |  |  |  |
| Chlorine is more reactive because it has more protons to attract the extra electron |  |  |  |
| Fluorine releases more energy than chlorine when it gains an electron because the electron is closer to the nucleus so it is more strongly attracted |  |  |  |